

Article

Thermal Analysis of Binary Mixtures of Imidazolium, Pyridinium, Pyrrolidinium, and Piperidinium Ionic Liquids

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Abstract: Ionic liquids (ILs) are being widely studied due to their unique properties, which make them potential candidates for conventional solvents. To study whether binary mixtures of pure ionic liquids provide a viable alternative to pure ionic liquids for different applications, in this work, the thermal analysis and molar heat capacities of five equimolar binary mixtures of ionic liquids based on imidazolium, pyridinium, pyrrolidinium, and piperidinium cations with dicyanamide, trifluoromethanesulfonate, and bis(trifluoromethylsulfonyl)imide anions have been performed. Furthermore, two pure ionic liquids based on piperidinium cation have been thermally characterized and the heat capacity of one of them has been measured. The determination and evaluation of both the transition temperatures and the molar heat capacities was carried out using differential scanning calorimetry (DSC). It was observed that the thermal behavior of the mixtures was completely different than the thermal behavior of the pure ionic liquids present, while the molar heat capacities of the binary mixtures were very similar to the value of the average of molar heat capacities of the two pure ionic liquids.

Keywords: ionic liquids; heat capacity; thermal behavior; heat capacity; DSC



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1. Introduction

In recent years, both the thermophysical properties and the application of ionic liquids (ILs) in different processes have been widely studied. This high interest in these compounds is mainly due to their special characteristics such as low vapor pressure and their high solvent power for a wide variety of chemical compounds, both polar and nonpolar. Moreover, they are liquid in a wide range of temperatures, which makes them more versatile.

One of the most important properties of ionic liquids is their so-called “tunability”, i.e., the possibility of synthesizing a suitable ionic liquid for a determined application. Within this approach, the possibility of mixing ionic liquids is a viable alternative, since mixed ILs can offer a significant improvement in the phase behavior properties and extend their applicability to other processes. There are many studies on the thermal analysis of pure ILs, however, there is a lack of studies related to heat capacities and transition temperatures of binary mixtures of ionic liquids in the literature [1–7].

In this context, with the purpose of calculating and comparing the thermal behavior and the molar heat capacities (C_p) of equimolar mixtures of two ionic liquids with the pure ionic liquids that form the mixture, the transition temperatures and the heat capacities of five equimolar mixtures of ionic liquids will be determined and evaluated using differential scanning calorimetry (DSC). The ionic liquids used in the mixtures were chosen so as to study the influence of the anion of the ionic liquid on the thermal properties of the mixtures. Therefore, the selected mixtures were the following: 1-butyl-3-methylimidazolium

bis(trifluoromethylsulfonyl)imide ((1,3) BMimNTf₂) + 1-butyl-3-methylimidazolium trifluoromethanesulfonate ((1,3) BMimTFO); 1-butyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide ((1,3) BMimNTf₂) + 1-butyl-3-methylimidazolium dicyanamide ((1,3) BMimDCA); 1-butyl-3-methylpyridinium bis(trifluoromethylsulfonyl)imide ((1,3) BmpyNTf₂) + 1-butyl-3-methylimidazolium trifluoromethanesulfonate ((1,3) BmpyTFO); 1-butyl-1-methylpyrrolidinium bis(trifluoromethylsulfonyl)imide ((1,1) BmpyrNTf₂) + 1-butyl-1-methylpyrrolidinium trifluoromethanesulfonate ((1,1) BmpyrTFO) and 1-methyl-1-propylpiperidinium bis(trifluoromethylsulfonyl)imide ((1,1) PMpipNTf₂) + 1-methyl-1-propylpiperidinium trifluoromethanesulfonate ((1,1) PMpipTFO). No results have been found for comparison purposes for these mixtures

Furthermore, the thermal behavior for two the pure ionic liquids 1-methyl-1-propylpiperidinium bis(trifluoromethylsulfonyl)imide and 1-methyl-1-propylpiperidinium trifluoromethanesulfonate and the molar heat capacity as a function of temperature for the 1-methyl-1-propylpiperidinium bis(trifluoromethylsulfonyl)imide were carried out using DSC. The transition temperatures of pure 1-methyl-1-propylpiperidinium bis(trifluoromethylsulfonyl)imide were compared with literature values [8,9]; for the other pure ionic liquid, no values have been found of heat capacities or transition temperatures in the literature.

The transition temperatures of the five equimolar mixtures and the two pure ionic liquids were carried out cooling from 393.15 K to 133.15 K and heating from 133.15 K to 393.15 K at a rate of 2 K·min^{−1}. The molar heat capacities for the mixtures were measured for the 258.15–348.15 K, and for the 1-methyl-1-propylpiperidinium bis(trifluoromethylsulfonyl)imide for the 293.15–393.15 K.

2. Results and Discussion

2.1. Pure Ionic Liquids

2.1.1. Thermal Behavior

Of the nine ionic liquids chosen to form the binary mixtures studied in this work, two of them (1-methyl-1-propylpiperidinium bis(trifluoromethylsulfonyl)imide and 1-methyl-1-propylpiperidinium trifluoromethanesulfonate) were thermally characterized by determining their transition temperatures and their Cp values as a function of temperature.

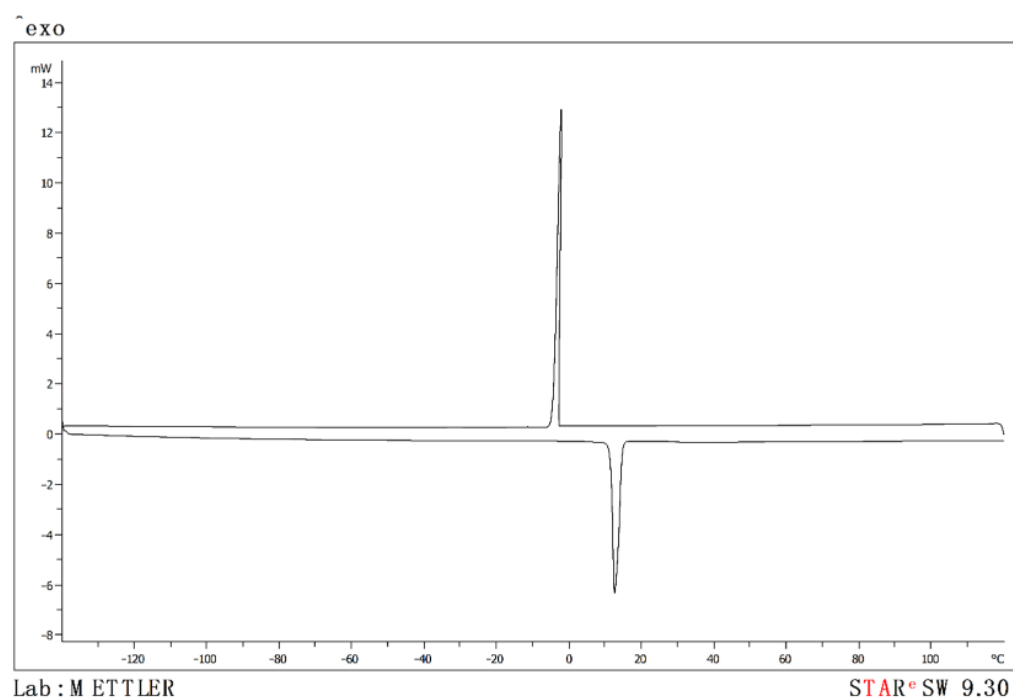
Table 1 presents the values of the transition temperatures determined with DSC for the pure ionic liquids ((1,1) PMpipNTf₂ and (1,1) PMpipTFO), together with the values found in the literature [8,9]. Figure 1 shows the respective thermograms.

Table 1. Results of the thermal analysis for the pure ionic liquids and comparison with the literature ^a.

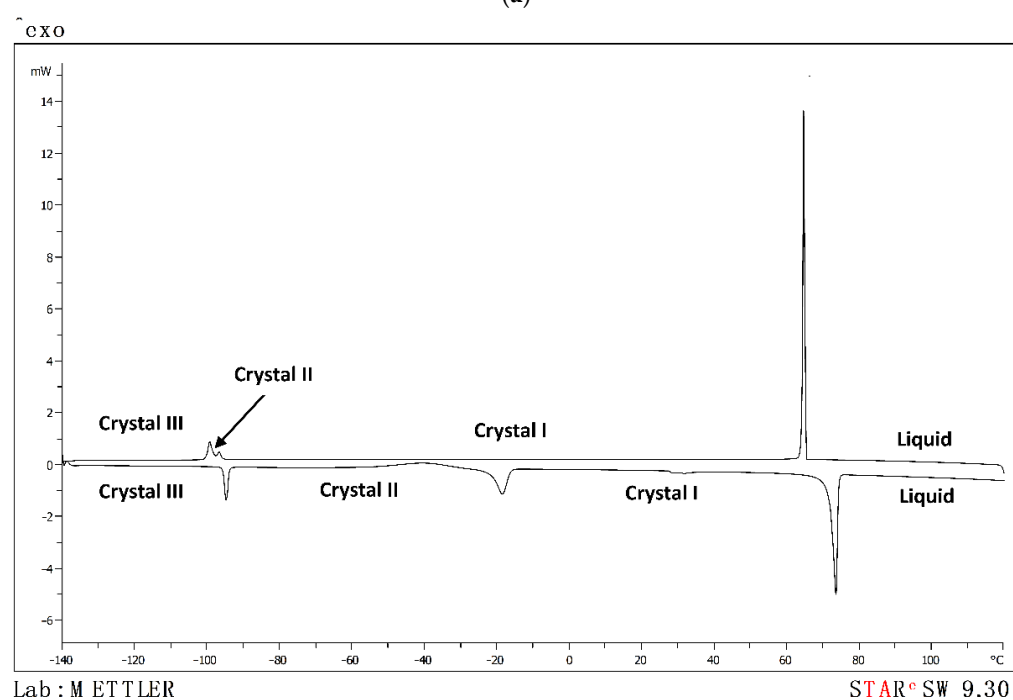
Ionic Liquid	T _{f(I-L)} / K	T _{f(I-II)} / K	T _{f(II-III)} / K	T _{ss(III-II)} / K	T _{ss(II-I)} / K	T _{m(I-L)} / K
(1,1) PMpipNTf ₂	273					283 (286) ^{b,c}
(1,1) PMpipTFO	173	183	323	173	258	323

^a Standard uncertainty $u(T) = \pm 1$ K; ^b from reference [8] at 10 K·min^{−1} rate; ^c from reference [9] at 10 K·min^{−1} rate.

As can be observed in Table 1 and Figure 1, the two piperidinium-based ionic liquids exhibit freezing and fusion points, therefore these ionic liquids are good crystal formers. The ionic liquid (1,1) PMpipNTf₂ only presents one freezing and fusion temperature (Figure 1a), while in the thermogram for the ionic liquid (1,1) PMpipTFO (Figure 1b) three peaks can be observed on heating and three peaks on cooling. The first peak on cooling (T_{f(I-L)}) corresponds to fusion, i.e., to the transition of the liquid state to the first solid state (crystal I). The second peak (T_{f(I-II)}) is the change of morphology of crystal I to crystal II, the second solid state. Lastly, the third peak (T_{f(II-III)}) corresponds to the conversion from crystal II to crystal III. Following the same logic, on heating, three peaks appear that could be related to melting points, however, the first peak (T_{m(III-II)}) corresponds to the change of state from crystal III to II, the second T_{m(II-I)} corresponds to the change of crystal II to crystal I and the last peak (T_{m(I-L)}) is the melting point of this ionic liquid. This behavior is characteristic of polymorphic materials [10].



(a)



(b)

Figure 1. Thermogram cooling the sample from 393.15 K to 133.15 K and heating from 393.15 K to 133.15 K at a rate of $2 \text{ K} \cdot \text{min}^{-1}$ for pure ionic liquids: (a) (1,1) PMpipNTf₂ and (b) (1,1) PMpipTFO.

When comparing the obtained results with data from the literature, also shown in Table 1, the experimental melting point for (1,1) PMpipNTf₂ is very similar to the bibliography, although the equipment and method used were different. For (1,1) PMpipTFO, no bibliography information was found.

2.1.2. Heat Capacities

Table S1 (available in Supporting Material) and Figure 2 present the experimental molar heat capacities for the pure ionic liquid evaluated in this work in the temperature range 308.15–363.15 K. A linear equation was used for the fitting of the molar heat capacity with temperature:

$$C_p = y_0 + aT \quad (1)$$

The fitting parameters (y_0 , and a) of Equation (1) and the correlation factor, R^2 , are shown in Table 2, and the graphical representation of C_p values as a function of temperature and the fitting can be seen in Figure 2. As can be observed from this figure, the heat capacity increases linearly with temperature, which is the expected trend and the common behavior on this property of ionic liquids [10–12].

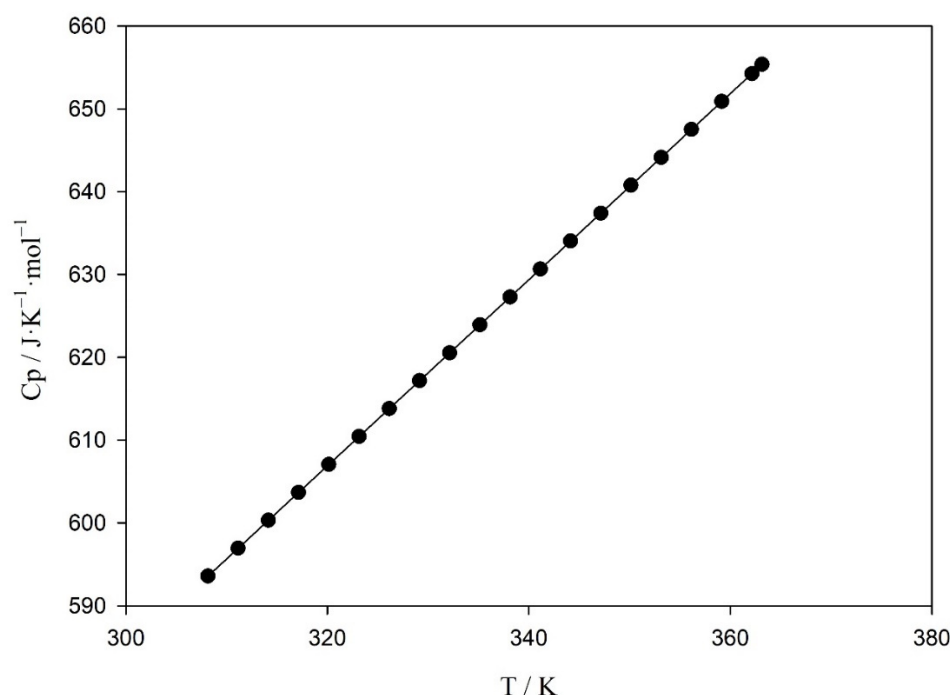


Figure 2. Experimental molar heat capacities, C_p , as a function of temperature for the (1,1) PMpipNTf₂, solid lines represent the fitting using Equation (1).

Table 2. Fitting parameters for the fitting ($C_p = y_0 + aT$) of the experimental molar heat capacities for PMpipNTf₂ ionic liquid.

IL	y_0	a	R^2
(1,1) PMpipNTf ₂	247.37	1.1235	0.9998

2.2. Mixtures of Ionic Liquids

2.2.1. Thermal Behavior

To determine and evaluate the transition temperatures in the DSC, five equimolar mixtures of different pairs of ionic liquids with imidazolium, pyridinium, pyrrolidinium, and piperidinium cations and dicyanamide, trifluoromethanesulfonate and bis(trifluorosulfonyl)imide anions have been prepared. The binary mixtures were: (1,3) BMimTFO + (1,3) BMimNTf₂; (1,3) BMimDCA + (1,3) BMimNTf₂; (1,3) BmpyTFO + (1,3) BmpyNTf₂; (1,1) BmpyrTFO + (1,1) BmpyrNTf₂ and (1,1) PMpipTFO + (1,1) PMpipNTf₂. They were selected with the aim of investigating the influence of the anion of the ionic liquid on the thermal behavior.

The values of the transition temperatures determined for the mixtures from thermal analysis and for the pure ionic liquids from previous works for the same group [10–12] can be observed in Table 3 in order to ease comparison. It can be concluded that the equimolar binary mixtures of ionic liquids present a totally different thermal behavior than the pure ionic liquids that constitute the mixture, since the mixtures evaluated in this work do not show crystallization in the range of temperatures studied and only present glass transition temperatures.

Table 3. Results of the thermal analysis for the pure ILs [10–12] and their mixtures ^a.

IL/Mixture of ILs	T _g / K	T _{f(L-I)} / K	T _{f(I-II)} / K	T _{f(II-III)} / K	T _{cc1} / K	T _{cc2} / K	T _{ss(III-II)} / K	T _{ss(II-I)} / K	T _{m(I-L)} / K
(1,3) BMimTFO ^b		253	258						289
(1,3) BMimDCA ^b		181				225		266	269
(1,3) BMimNTf ₂ ^c	184								
(1,3) BMimTFO + (1,3) BMimNTf ₂	178								
(1,3) BMimDCA + (1,3) BMimNTf ₂	181								
(1,3) BMpyTFO ^b		255	256						
(1,3) BMpyNTf ₂ ^b	188								
(1,3) BMpyTFO + (1,3) BMpyNTf ₂	188								
(1,1) BMpyrTFO ^d		245						238	267
(1,1) BMpyrNTf ₂ ^d	183				208			244	253
(1,1) BMpyrTFO + (1,1) BMpyrNTf ₂	183								
(1,1) PMpipTFO ^e		173	183	333			173	258	333
(1,1) PMpipNTf ₂ ^e		273							283
(1,1) PMpipTFO + (1,1) PMpipNTf ₂	198								

^a Standard uncertainty $u(T) = \pm 1$ K; ^b from reference [11]; ^c from reference [10]; ^d from reference [12]; ^e this work.

Figure 3 shows, for example, the thermogram for mixture (1,1) PMpipTFO + (1,1) PMpipNTf₂. The remaining thermograms are available in the Supporting Material (Figure S1).

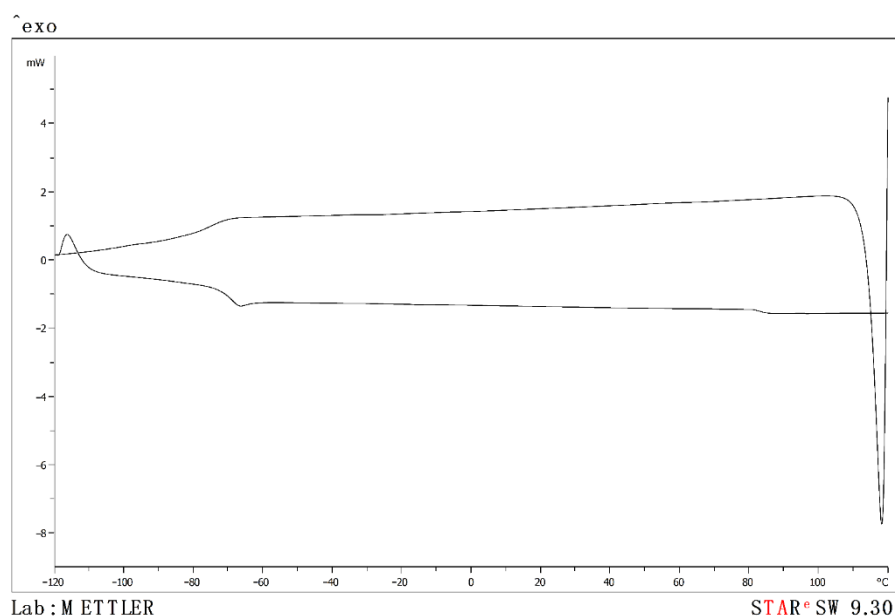


Figure 3. Thermogram cooling the sample from 393.15 K to 133.15 K and heating from 393.15 K to 133.15 K at a rate of 2 K·min^{−1} for the equimolar binary mixture (1,1) PMpipTFO + (1,1) PMpipNTf₂.

2.2.2. Heat Capacities

The molar heat capacity of the five equimolar binary mixtures of ionic liquids has been determined by DSC with the method of sapphire in the range of 298.15–348.15 K. In Table S2 (Supporting Material) and Figure 4, the experimental molar heat capacities for the mixtures of ionic liquids in the temperature range 293.15–333.15 K are presented.

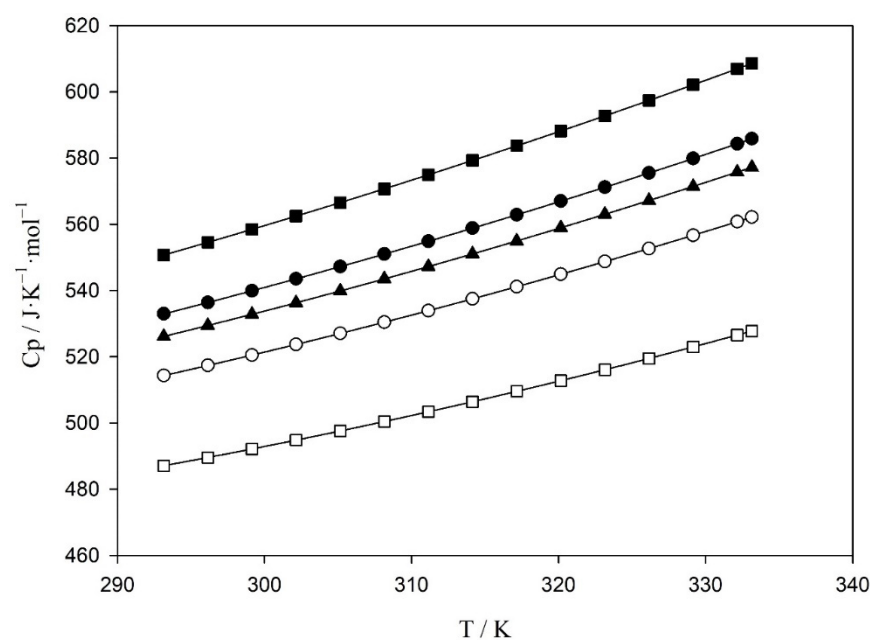


Figure 4. Experimental molar heat capacities, C_p , as a function of temperature for the equimolar binary mixtures: () (1,3) BMimTFO + (1,3) BMimNTf₂; () (1,3) BMimDCA + (1,3) BMimNTf₂; (▲) (1,3) BMpyTFO + (1,3) BMpyNTf₂; (●) (1,1) BMpyrTFO + (1,1) BMpyrNTf₂ and (■) (1,1) PMpipTFO + (1,1) PMpipNTf₂, solid lines represent the fitting using the Equation (2).

Table 4 shows the results at 323 K and the C_p of the pure ionic liquids at the same temperature [10–12] for comparison purposes. The five binary mixtures showed a value of C_p which is intermediate of the values of the individual ionic liquids, therefore, these C_p values reveal that the C_p of the two pure ionic liquids mixtures exhibits a practically ideal behavior, this behavior was also observed and reported in the bibliography [8] for mixtures of other ionic liquids. Furthermore, it can be seen that the piperidinium-based ionic liquid present higher heat capacity than their analogous pyrrolidinium, pyridinium, and imidazolium-based ionic liquids, in agreement with previous publications on the topic [6].

Table 4. Experimental heat capacities, C_p , for the pure ILs [10–12] and their mixtures at $T = 323.15$ K ^a.

Ionic Liquid or Mixture	$C_p / \text{J} \cdot \text{K}^{-1} \cdot \text{mol}^{-1}$
(1,3) BMimTFO ^b	456
(1,3) BMimDCA ^b	387
(1,3) BMimNTf ₂ ^c	551
(1,3) BMimTFO + (1,3) BMimNTf ₂	549
(1,3) BMimDCA + (1,3) BMimNTf ₂	516
(1,3) BMpyNTf ₂ ^b	572
(1,3) BMpyTFO ^b	461
(1,3) BMpyTFO + (1,3) BMpyNTf ₂	563
(1,1) BMpyrNTf ₂ ^c	590
(1,1) BMpyrTFO ^c	465
(1,1) BMpyrTFO + (1,1) BMpyrNTf ₂	571
(1,1) PMpipNTf ₂ ^d	610
(1,1) PMpipTFO ^d	n.a. ^e
(1,1) PMpipTFO + (1,1) PMpipNTf ₂	593

^a Standard uncertainty $u(T) = \pm 5\%$; ^b from reference [11]; ^c from reference [10]; ^d from reference [12]; ^e This work; n.a.—not available.

The fitting of the molar heat capacity of the equimolar mixtures with temperature was carried out using a polynomial equation:

$$C_p = y_0 + aT + bT^2 \quad (2)$$

Table 5 shows the fitting parameters (y_0 , a , and b) and the correlation factor, R^2 . Moreover, in Figure 4, the fitting is plotted. As can be observed in this figure, the dependence of the C_p of the equimolar mixtures of two ionic liquids with the temperature is practically linear.

Table 5. Fitting parameters for the fitting ($C_p = y_0 + aT + bT^2$) of the experimental molar heat capacities for equimolar binary mixtures.

Binary Mixtures	y_0	a	b	R^2
(1,3) BMimTFO + (1,3) BMimNTf ₂	640.2	1.86	0.0049	0.9999
(1,3) BMimDCA + (1,3) BMimNTf ₂	687.0	2.18	0.0051	0.9999
	617.5			
(1,3) BMpyTFO + (1,3) BMpy NTf ₂	605.8	1.71	0.0048	0.9999
	570.1			
(1,1) BMpyrTFO + (1,1) BMpyrNTf ₂	605.8	1.63	0.0047	0.9999
(1,1) PMpipTFO + (1,1) PMpipNTf ₂	570.1	1.40	0.0045	0.9999

3. Materials and Methods

3.1. Materials

The name, abbreviation used in this work, purity (in mole fraction), and water content of the nine ionic liquids used in this study are specified in Table 6.

Table 6. Name, abbreviation, purity, and water content of pure ionic liquids.

Name	bbreviation	Purity ^a (Mole Fraction)	Water Content ^b (ppm)
1-butyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide	(1,3) BMimNTf ₂	98%	<100
1-butyl-3-methylimidazolium trifluoromethanesulfonate	(1,3) BMimTFO	98%	<300
1-butyl-3-methylimidazolium dicyanamide	(1,3) BMimDCA	99%	<300
1-butyl-3-methylpyridinium bis(trifluoromethylsulfonyl)imide	(1,3) BMpyNTf ₂	98%	<300
1-butyl-3-methylpyridinium trifluoromethanesulfonate	(1,3) BMpyTFO	98%	<300
1-butyl-1-methylpyrrolidinium bis(trifluoromethylsulfonyl)imide	(1,1) BMpyrNTf ₂	99%	<100
1-butyl-1-methylpyrrolidinium trifluoromethanesulfonate	(1,1) BMpyrTFO	99%	<100
1-methyl-1-propylpyrrolidinium bis(trifluoromethylsulfonyl)imide	(1,1) PMpipNTf ₂	99%	<200
1-methyl-1-propylpyrrolidinium trifluoromethanesulfonate	(1,1) PMpipTFO	99%	<200

^a Mole fraction purity and analysis method given by the company; ^b standard uncertainty u (water content) = $\pm 5\%$.

The nine ionic liquids were purchased from IoLITEC (Heilbronn, Germany) and the water content given in Table 1 corresponds to the amount of water determined after purification (vacuum drying at 343 K and 0.2 Pa for at least 48 h) using a coulometric Karl Fischer titrator, model C20, using Hydranal-Coulomat CG and Hydranal-Coulomat AG as cathodic and anodic titrant. The measurements of water content were carried out prior to the determination of the molar heat capacities and transition temperatures.

3.2. Methods

The determination of the thermal behavior and the heat capacities of the pure ionic liquids and their binary mixtures were carried out using a Mettler-Toledo differential scanning calorimeter (DSC) (Mettler-Toledo, OH, USA), model DSC822e, and evaluated with the Mettler-Toledo STAR^e software version 9.30 (Mettler-Toledo, OH, USA).

The determination of transition temperatures was performed according to the following procedure: the samples of pure ionic liquids or their equimolar mixtures were cooled from 393.15 K to 133.15 K with a relatively slow heating rate of 2 K·min^{−1} and, afterwards, they were heated from 133.15 K to 393.15 K at the same rate. Therefore, before performing the thermal analysis, the ionic liquids were dried inside the DSC by heating at T = 393.15 K for 30 min.

To allow a good comparison with the literature of the temperatures determined in this work, it is important to understand that: the melting temperature (T_m) was taken as the onset of an endothermic peak on heating; the freezing temperature (T_f) as the onset of an exothermic peak on cooling; the glass transition temperature (T_g) as the midpoint of a small heat capacity change on heating from the amorphous glass state to a liquid state; the cold crystallization temperature (T_{cc}) as the onset of an exothermic peak on heating from a subcooled liquid state to a crystalline solid state; the solid-solid transition (T_{ss}) as the onset of an exothermic or endothermic peak on heating from a crystalline solid state.

For the thermal analyses, 40 L aluminum pans hermetically sealed with a pinhole at the top were used for the samples and a similar empty pan (also with a pinhole) was used as a reference in the furnace. The masses of the samples were between 4 and 8 mg. All the samples were weighed using a Mettler-Toledo AX-205 Delta Range balance with an uncertainty in the measurement of $\pm 3 \times 10^{-4}$. The calibration for temperature and heat flow was performed using the following pure substances: zinc, indium, water, and heptane. The tau lag calibration was carried out at the 20, 15, 10, 2.5, 10, and 20 K·min^{−1} rates. The standard uncertainty of the measurement of the temperature in the thermal analysis was ± 1 K.

The sapphire method was chosen for the determination of the value of the molar heat capacity as a function of the temperature of the pure ionic liquids and their binary mixtures. The sapphire method consists of an initial isothermal segment for 15 min followed by a dynamic period at 20 K·min^{−1} and finally an isothermal segment for 15 min; 100 L aluminum pans hermetically sealed with a pinhole at the top were used. The uncertainty of the measurement of the molar heat capacity was $\pm 5\%$. This procedure for thermal analysis and measurement of heat capacities and the calculation of the standard uncertainties was described thoroughly in a previous publication [13].

Heat capacities were determined from different temperature ranges in the function of the melting point of the pure ionic liquid since heat capacity was only measured for the liquid range of the ionic liquids, therefore, the volume variation in the temperatures range studied can be considered negligible. For the equimolar binary mixture of two ionic liquids, the range was 258.15–348.15 K, since all mixtures studied are liquid at ambient temperature and pressure. Only data from 293.15 K to 333.15 K are presented, since, at the temperature range limits, the samples were not perfectly thermally conditioned. Therefore, the molar heat capacity for the ionic liquid 1-propyl-1-methylpiperidinium bis(trifluoromethylsulfonyl)imide was determined from 293.15–393.15 K, since this ionic liquid has a melting point at 283 K. The C_p measurements were carried out in the liquid range of the ILs, but only data from 308.15–363.15 K are presented. The C_p for the pure 1-propyl-1-methylpiperidinium trifluoromethanesulfonate ionic liquid was not determined, since this ionic liquid is solid up to a temperature of 333 K.

4. Conclusions

In this work, the thermal behavior and the molar heat capacities for 5 equimolar binary mixtures of ionic liquids were determined and evaluated by DSC. The results obtained in this work for the mixtures were compared with the thermal behavior and the molar heat

capacities of the respective pure ionic liquids previously published by the group. Moreover, the transition temperatures for two pure piperidinium-based ionic liquids and the heat capacities for one pure piperidinium-based ionic liquid were measured in an assay.

The results show that mixing different ionic liquids leads to a new material with improved characteristics when compared to the starting components. The thermal behavior of the binary mixtures of the ionic liquids (which present melting points) demonstrated that the mixtures do not have melting points, but only present glass transition temperatures.

It was shown that the C_p values of the binary mixtures of ionic liquids are in the interval between the values of C_p for the pure ionic liquids present in the mixture. Furthermore, the trend found for the value of the heat capacities of the pure ionic liquid was as follows: piperidinium > pyrrolidinium > pyridinium > imidazolium-based ionic liquids.

Supplementary Materials: The following are available online, Table S1. Experimental molar heat capacities, C_p , as function of temperature for the (1,1) PMpipNTf₂ ionic liquid. Table S2. Experimental molar heat capacities, C_p , as function of temperature for equimolar binary mixtures. Figure S1. Thermogram cooling the sample from 393.15 K to 133.15 K and heating from 393.15 K to 133.15 K at a rate of 2 K·min^{−1} for (a) (1,3) BMimTFO + (1,3) BMimNTf₂; (b) (1,3) BMimDCA + (1,3) BMimNTf₂; (c) (1,3) BMpyTFO + (1,3) BMpyNTf₂ and (d) (1,1) BMpyrTFO + (1,1) BMpyrNTf₂.

Author Contributions: Conceptualization, E.G. and Á.D.; methodology, E.G. and P.V.; validation, E.G. and P.V.; formal analysis, E.G., E.A.M. and Á.D.; investigation, E.G. and P.V.; data curation, E.G. and Á.D.; writing—original draft preparation, E.G. and Á.D.; writing—review and editing, E.G., E.A.M. and Á.D.; supervision, E.A.M. and Á.D.; funding acquisition, E.A.M. All authors have read and agreed to the published version of the manuscript.

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Conflicts of Interest: The authors declare no conflict of interest.

Sample availability: Samples of the compounds are not available from the authors.

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